Atmospheric Chemistry

- Structure and Circulation
- Gas Composition
- Photochemistry
- Case Study: Ozone

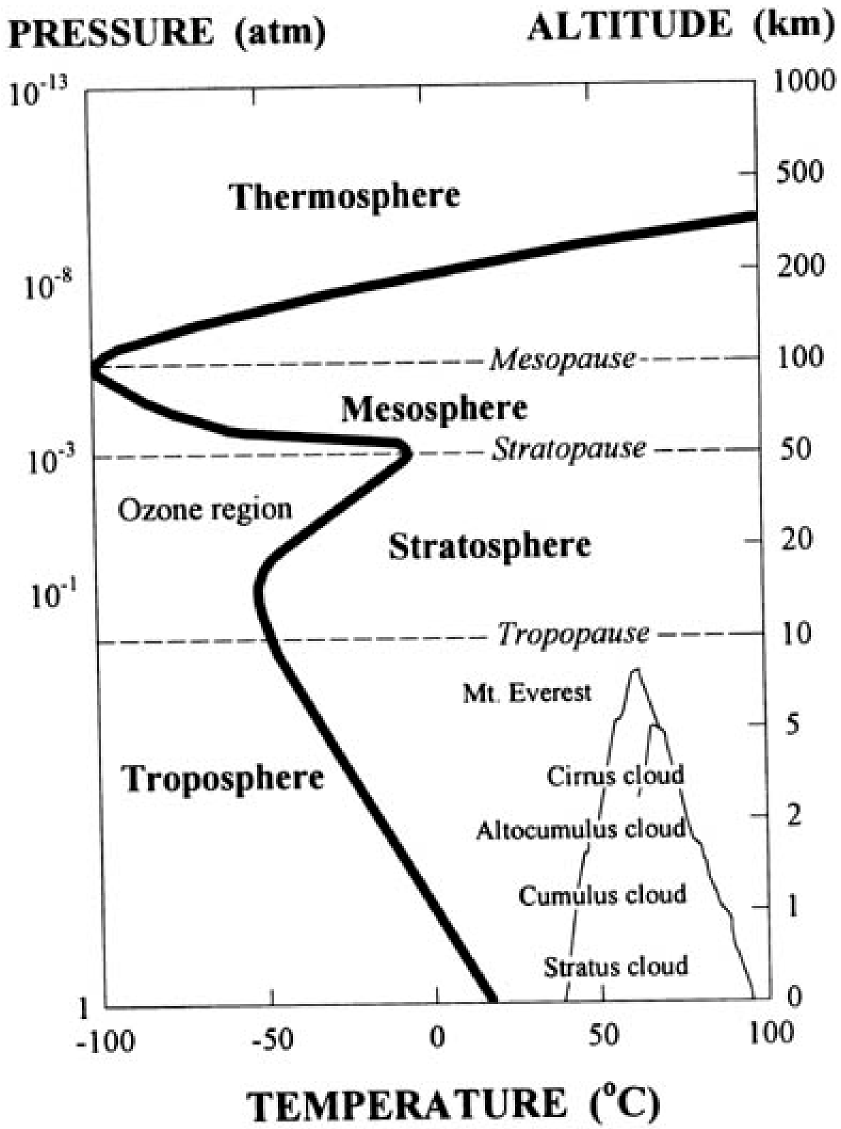
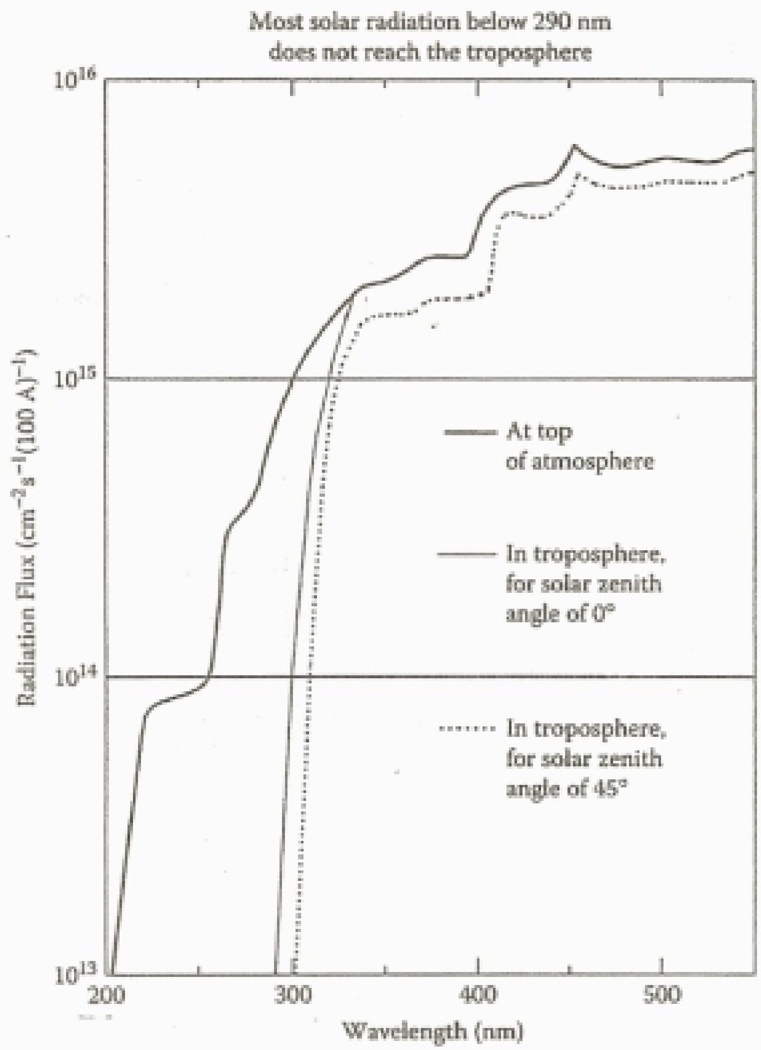


FIGURE 5.1. Changes in temperature in the layers of the atmosphere.



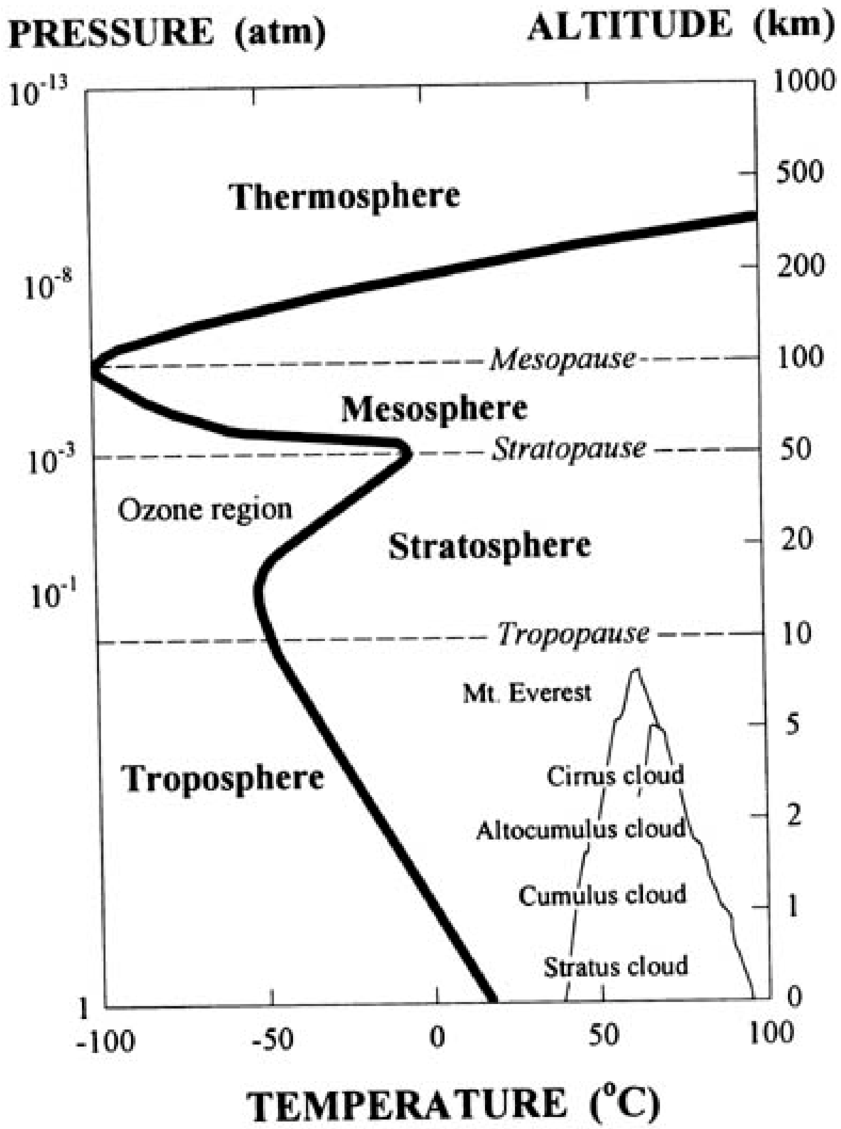


FIGURE 5.1. Changes in temperature in the layers of the atmosphere.

Thermosphere (Above 100 km)

N_2 and O_2 absorb light below 240 nm

Stratosphere (10 to 50 km)

O_3 absorbs light between 240 and 300 nm

$$O_2 + h\nu \rightarrow 2\,O^{\cdot}$$
$$O_2 + O^{\cdot} + M \rightarrow O_3 + M$$

$$O_3 + h\nu \rightarrow O_2 + O(^1D)$$

Troposphere (below 10 km)

no light available to break O-O bonds
> 300 nm of 120 kcal bond strength

Oxidizer is the $OH^{\cdot}$ radical not O_3 or H_2O_2

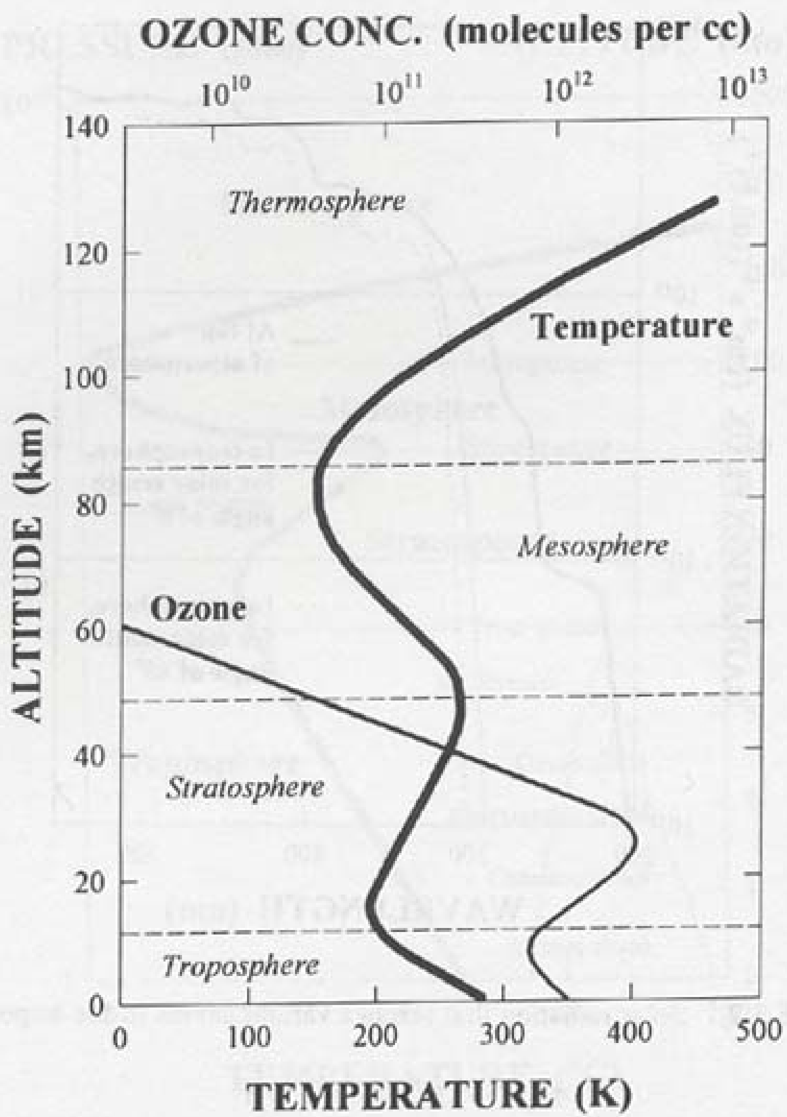
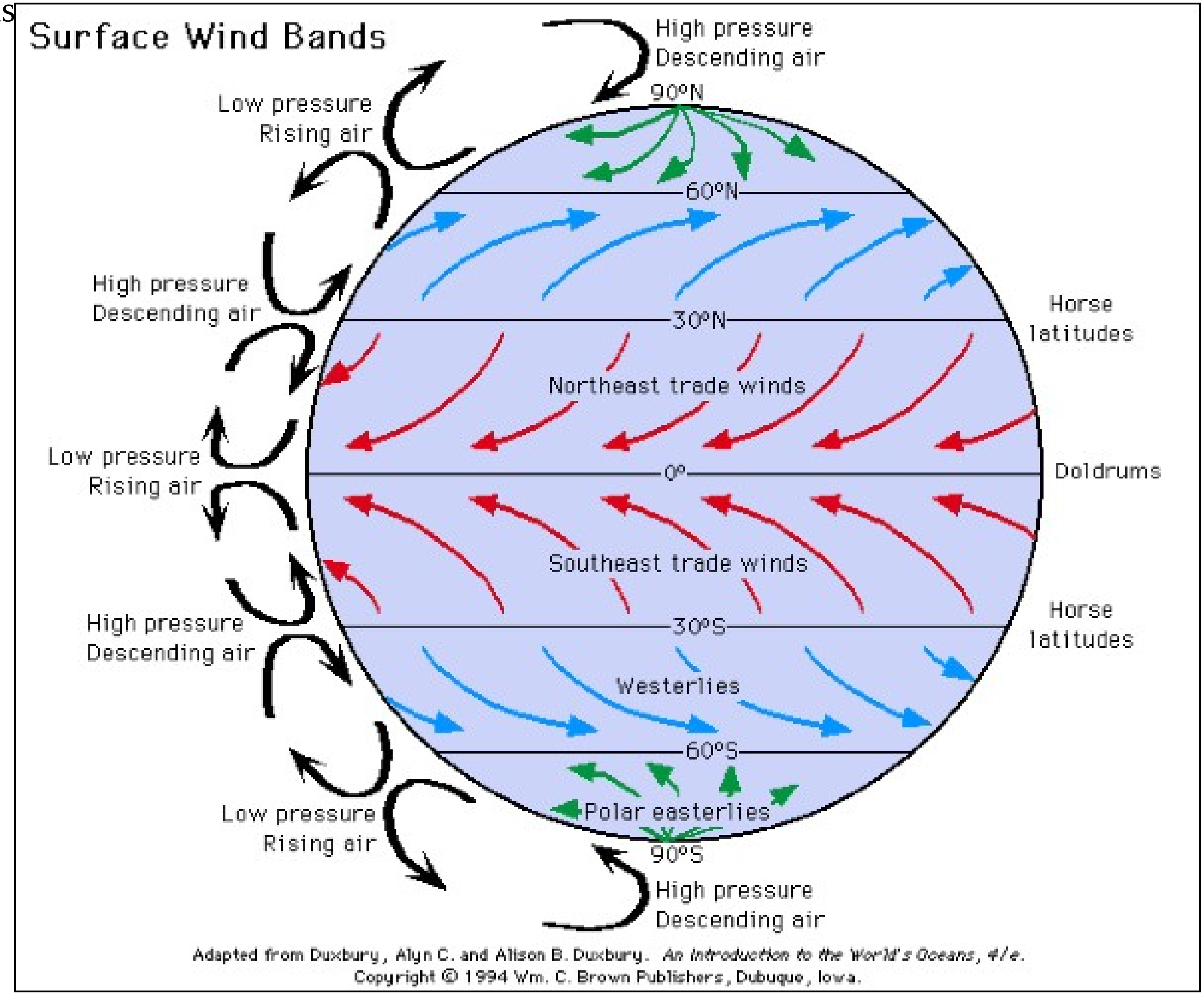


FIGURE 5.3. Concentration of ozone in various layers of the atmosphere.

Winds



Surface Wind Bands
High pressure
Descending air
90°N
Low pressure
Rising air
60°N
High pressure
Descending air
30°N
Horse
latitudes
Northeast trade winds
Low pressure
Rising air
0°
Doldrums
Southeast trade winds
High pressure
Descending air
30°S
Horse
latitudes
Westerlies
60°S
Low pressure
Rising air
Polar easterlies
90°S
High pressure
Descending air
Adapted from Duxbury, Alyn C. and Alison B. Duxbury. An Introduction to the World's Oceans, 4/e.
Copyright © 1994 Wm. C. Brown Publishers, Dubuque, Iowa.

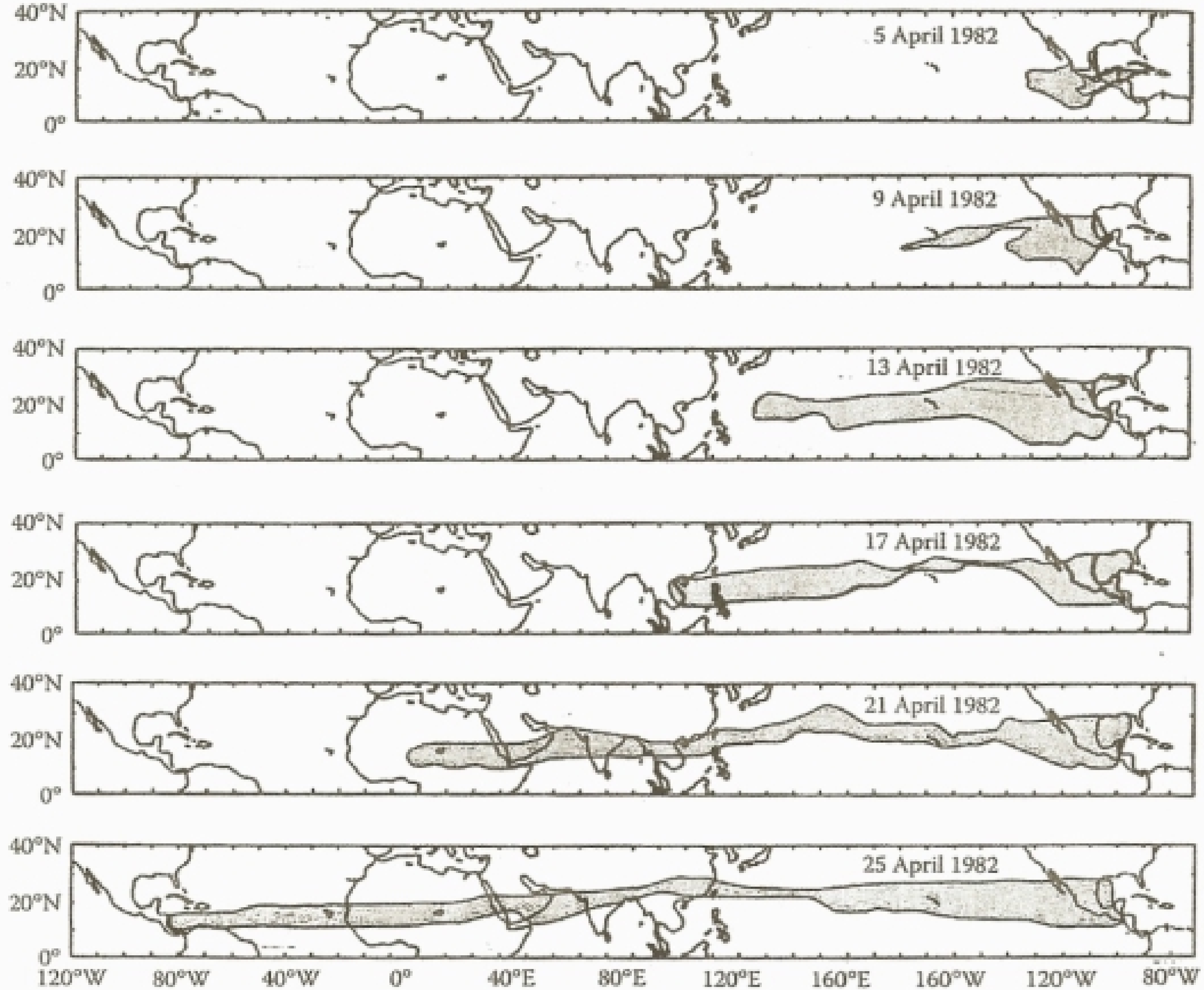


FIGURE 5.9
The movement of dust and particles in the northern hemisphere after the El Chichon Volcano eruption.

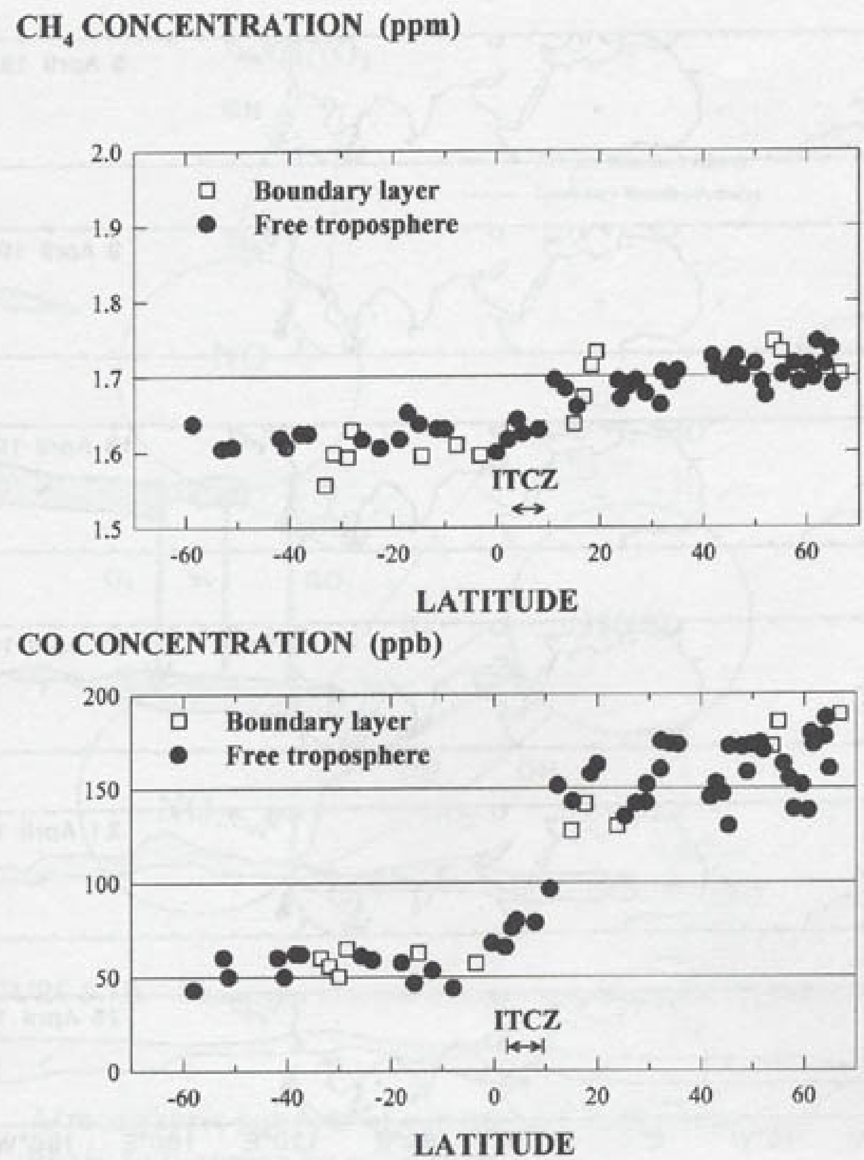


FIGURE 5.8. Distribution of methane (top) and carbon monoxide (bottom) between the hemispheres.

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TABLE 5.1
Abundance of the Major Conservative Atmospheric Gases[a]

Gas	Mole fraction in dry air (X_i)
N_2	0.78084 ± 0.00004
O_2	0.20946 ± 0.00002
Ar	$(9.34 \pm 0.01) \times 10^{-3}$
CO_2	$(3.5 \pm 0.1) \times 10^{-4}$
Ne	$(1.818 \pm 0.004) \times 10^{-5}$
He	$(5.24 \pm 0.004) \times 10^{-6}$
Kr	$(1.14 \pm 0.01) \times 10^{-6}$
Xe	$(8.7 \pm 0.1) \times 10^{-8}$

[a] Kester (1975).

TABLE 6.2
Isotopic Abundance of Atmospheric Gases[a]

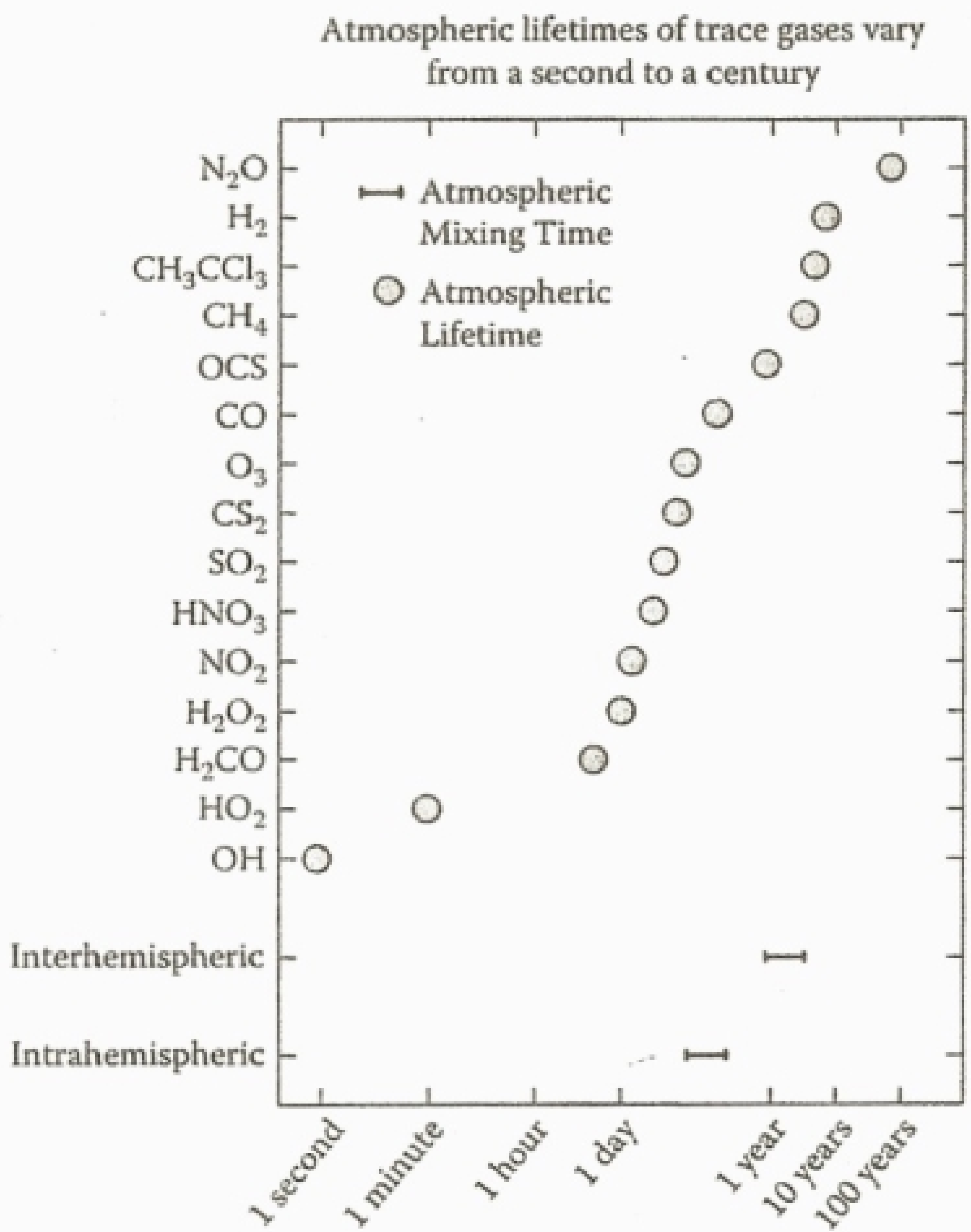
Element	Mass number	Mole %	Element	Mass number	Mole %
H (in H_2O)	1	99.98	Kr	78	0.354
H (in H_2O)	2	0.02	Kr	80	2.27
He	3	1.1×10^{-4}	Kr	82	11.56
He	4	100.0	Kr	83	11.55
C (in CO_2)	12	98.9	Kr	84	56.90
C (in CO_2)	13	1.1	Kr	86	17.37
C (in CO_2)	14	9.5×10^{-13}			
N	14	99.62	Xe	124	0.096
N	15	0.38	Xe	126	0.090
O	16	99.757	Xe	128	1.919
O	17	0.039	Xe	129	26.44
O	18	0.204	Xe	130	4.08
Ne	20	90.92	Xe	131	21.18
Ne	21	0.257	Xe	132	26.89
Ne	22	8.82	Xe	134	10.44
Ar	36	0.337	Xe	136	8.87
Ar	38	0.063			
Ar	40	99.600			

[a] Kester (1975).

TABLE 5.2

The Composition of Minor Gases in the Atmosphere

Species	X_i actual	Reliability	Source	Sink
CH_4	1.7×10^{-6}	High	Biog.	Photochem.
CO	$0.5–2 \times 10^{-7}$	Fair	Photo., anthr.	Photochem.
O_3	5×10^{-8} (clean)	Fair	Photo	Photochem.
	4×10^{-7} (polluted)			
	10^{-7} to 6×10^{-6} (stratosphere)			
$NO + NO_2$	$10^{-12}–10^{-8}$	Low	Lightn., anthr. photo.	Photochem.
HNO_3	$10^{-11}–10^{-9}$	Low	Photo.	Rainout
NH_3	$10^{-10}–10^{-9}$	Low	Biog.	Photo., rainout
N_2O	3×10^{-7}	High	Biog.	Photo.
H_2	5×10^{-7}	High	Biog., photo.	Photo.
OH	$10^{-15}–10^{-12}$	Very low	Photo.	Photo.
HO_2	$10^{-11}–10^{-13}$	Very low	Photo.	Photo.
H_2O_2	$10^{-10}–10^{-18}$	Very low	Photo.	Rainout
H_2CO	$10^{-10}–10^{-9}$	Low	Photo.	Photo.
SO_2	$10^{-11}–10^{-10}$	Fair	Anth., photo.,	Photo., volcanic
CS_2	$10^{-11}–10^{-10}$	Low	Anthr., biol.,	Photo.
OCS	5×10^{-10}	Fair	Anthr., biol., photo.	Photo.
CH_3CCl_3	$0.7–2 \times 10^{-10}$	Fair	Anthropogenic	Photo.


Atmospheric lifetimes of trace gases vary
from a second to a century
Atmospheric
Mixing Time
Atmospheric
Lifetime
N2O
H2
CH3CCl3
CH4
OCS
CO
O3
CS2
SO2
HNO3
NO2
H2O2
H2CO
HO2
OH
Interhemispheric
Intrahemispheric
1 second
1 minute
1 hour
1 day
1 year
10 years
100 years

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Production of OH Radicals

$$O_3 + h\nu\ (315\text{-}1200\ nm) \longrightarrow O(^3P) + O_2$$

$$O(^3P) + O_2 \longrightarrow O_3$$

$$O_3 + h\nu\ (< 315\ nm) \longrightarrow O(^1D) + O_2$$

$$O(^1D) + H_2O \longrightarrow 2\ OH\cdot$$

Removal Of OH Radicals

$$CO + OH\cdot \longrightarrow CO_2 + H\cdot$$

$$CH_4 + OH\cdot \longrightarrow \cdot CH_3 + H_2O$$

$$H + O_2 \longrightarrow HO_2$$

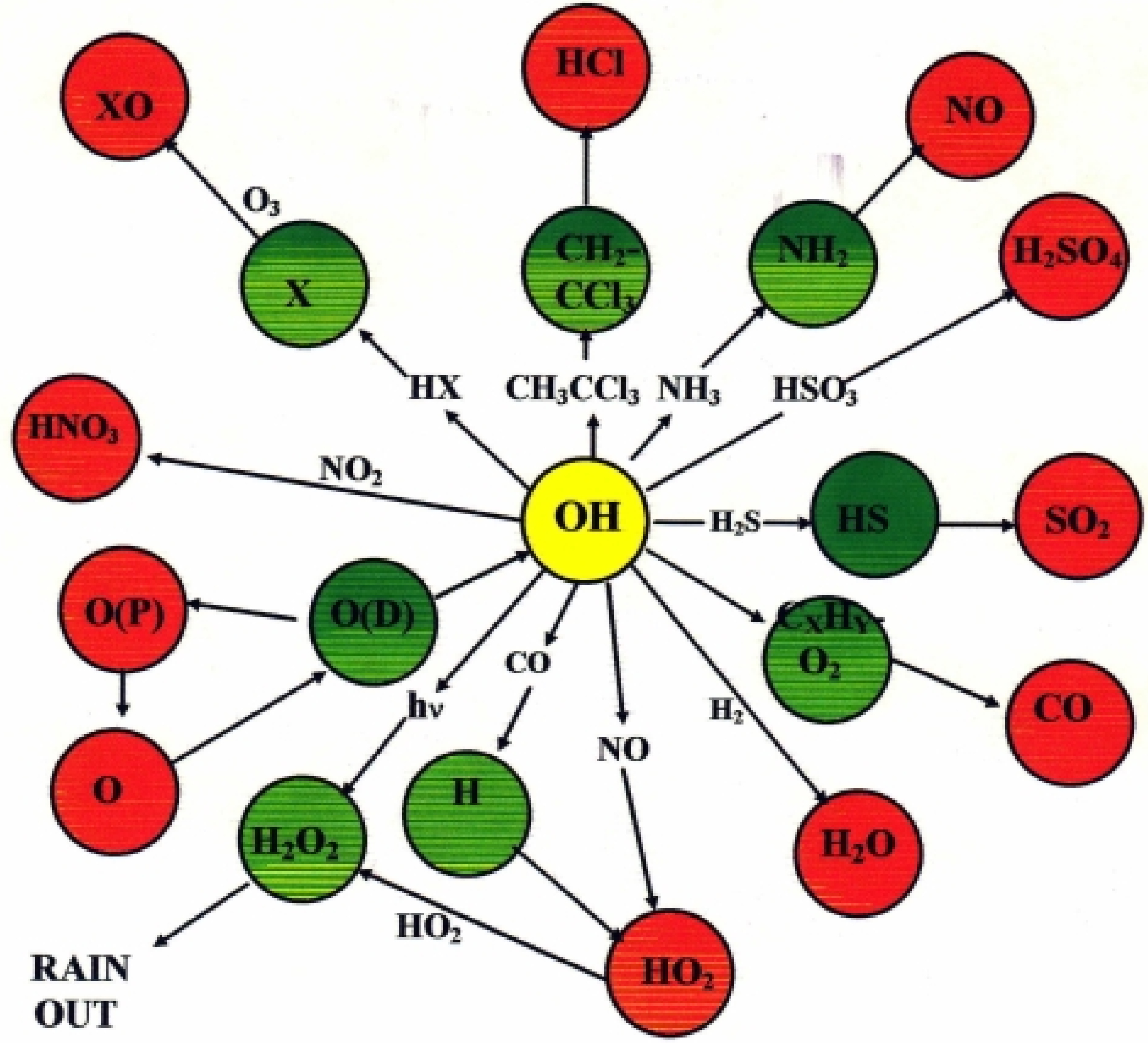
$$HO_2 + NO \longrightarrow NO_2 + OH\cdot$$

$$HO_2 + O_3 \longrightarrow 2\ O_2 + OH\cdot$$

$$HO_2 + OH\cdot \longrightarrow H_2O + O_2$$

$$HO_2 + HO_2 \longrightarrow H_2O_2 + O_2$$ Chain Termination

Photochemistry of the OH Radical



SUMMARY OF IMPORTANT NITROGEN OXIDE REACTIONS

(Species NO, NO_2, HNO_3)

Gas Phase

$N_2 + O_2 \longrightarrow 2\ NO$ High Temperatures

$NO + HO_2\ (RO_2) \longrightarrow NO_2 + OH^{\bullet}\ (RO^{\bullet})$

$NO_2 + h\nu \longrightarrow NO + O^{\bullet}$

$O_2 + O^{\bullet} + M \rightarrow O_3 + M$

$NO_2 + OH^{\bullet} \longrightarrow HNO_3$

Aqueous Phase

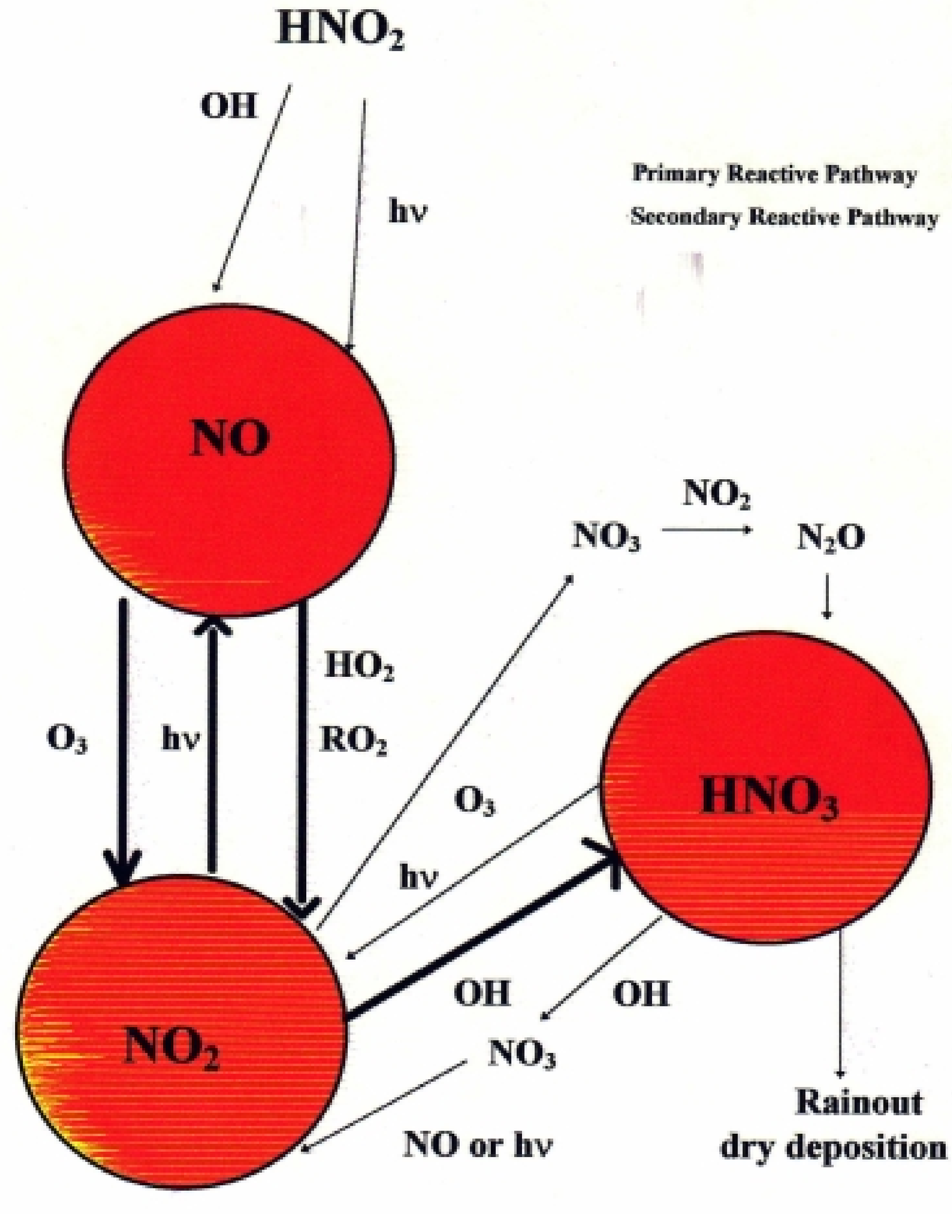
$NO_2 + O_3 \longrightarrow NO_3 + O_2$

$NO_3 + NO_2 \longrightarrow N_2O_5$

$N_2O_5 + H_2O \longrightarrow 2\ HNO_3$

$2\ NO_2 + O_3 + H_2O \longrightarrow 2\ HNO_3 + O_2$

Most important source of O_3 and other oxidants in the Troposphere.
Hydrocarbons are needed to produce HO_2.


HNO2
OH
hν
Primary Reactive Pathway
Secondary Reactive Pathway
NO
NO3
NO2
N2O
HO2
O3
hν
RO2
O3
HNO3
hν
OH
OH
NO3
NO2
NO or hν
Rainout
dry deposition

SULFUR OXIDE CHEMISTRY

Gas

$SO_2 + OH\cdot \longrightarrow HOSO_2$

$HOSO_2 + O_2 \longrightarrow SO_3 + HO_2$

$SO_3 + H_2O \longrightarrow H_2SO_4$

Aqueous

$SO_2 + H_2O \longrightarrow HSO_3^- + H^+ \longrightarrow SO_3^{-2} + H^+$

$HSO_3\text{-} + H_2O_2 \longrightarrow SO_4^{-2} + H^+ + H_2O$

$SO_3^{-2} + O_3 \longrightarrow SO_4^{-2} + O_2$

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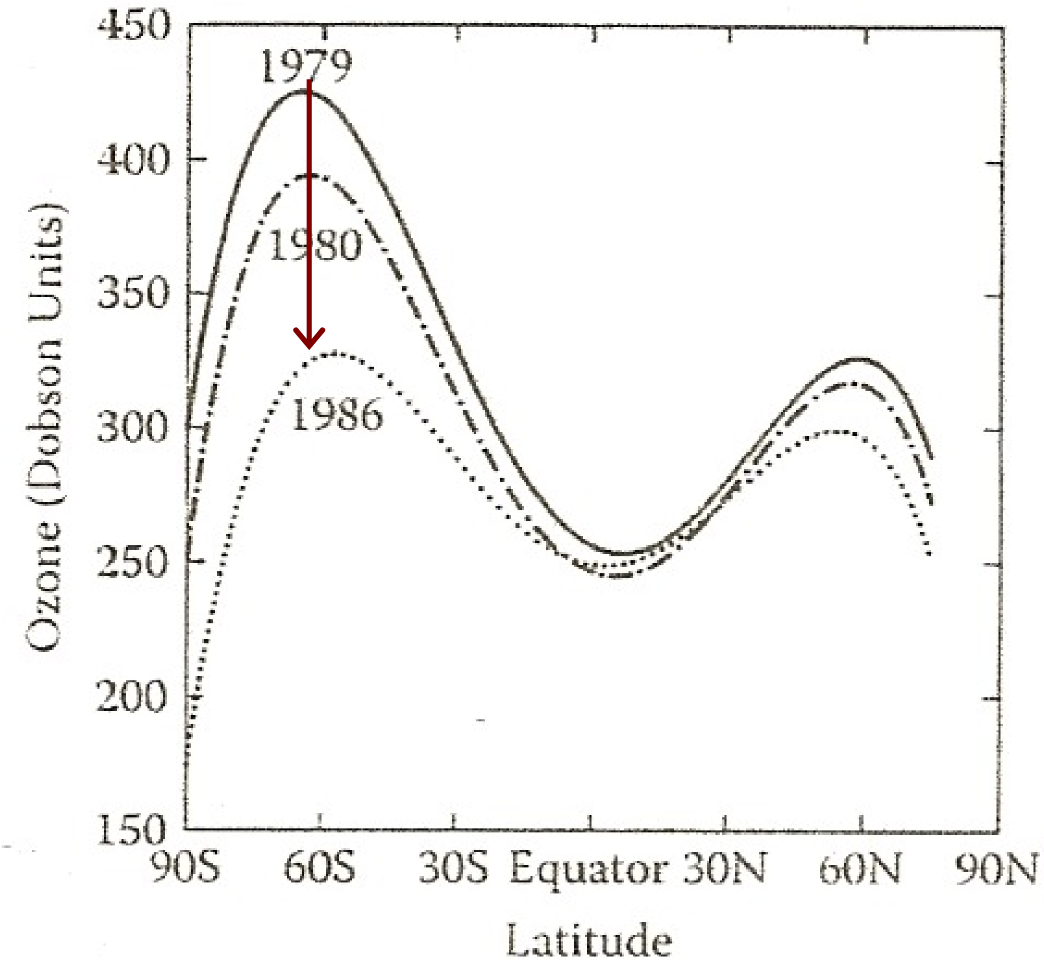
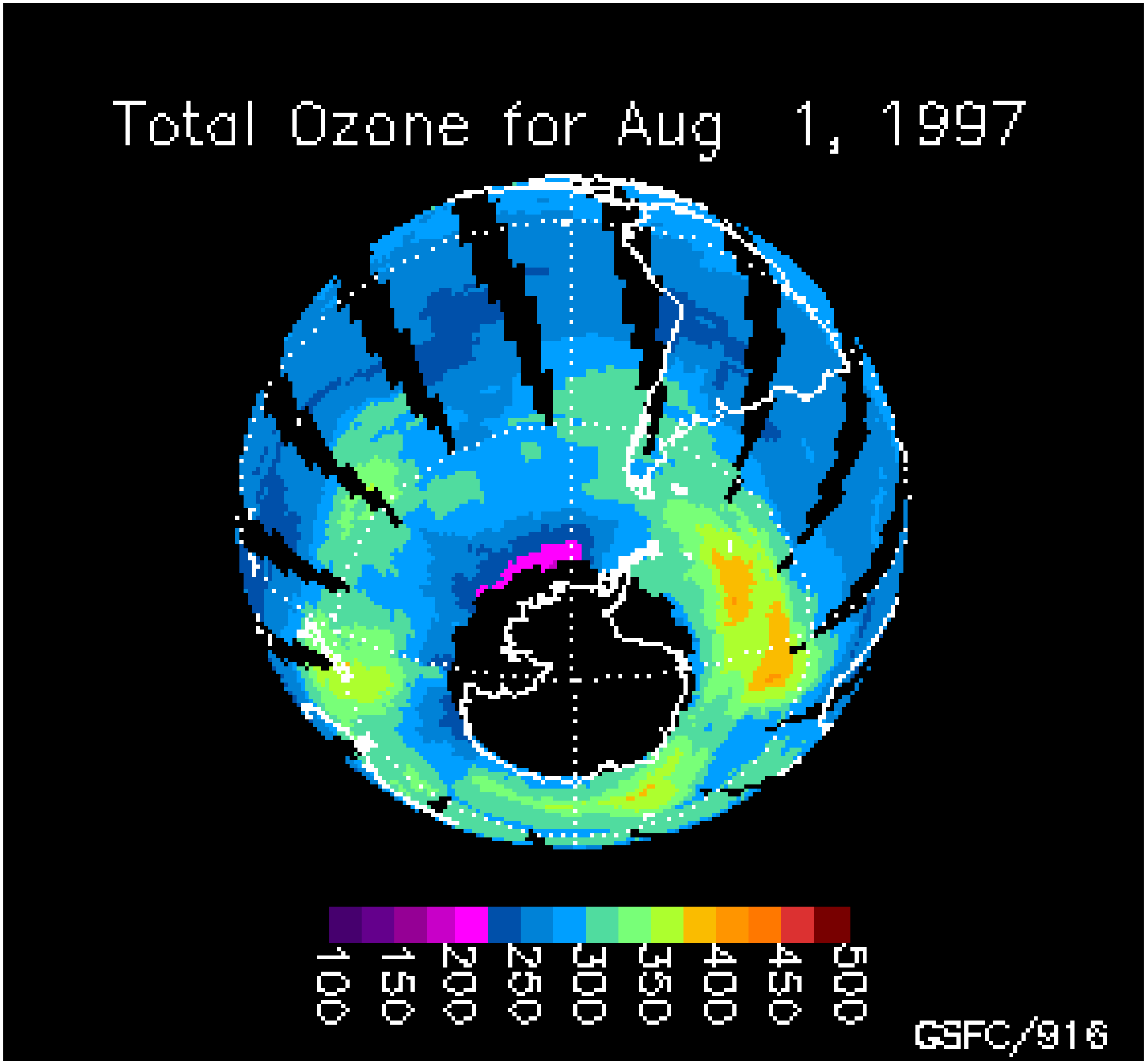


FIGURE 5.37
The levels of ozone as a function of latitude.



Total Ozone for Aug 1, 1997
100
150
200
250
300
350
400
450
500
GSFC/916

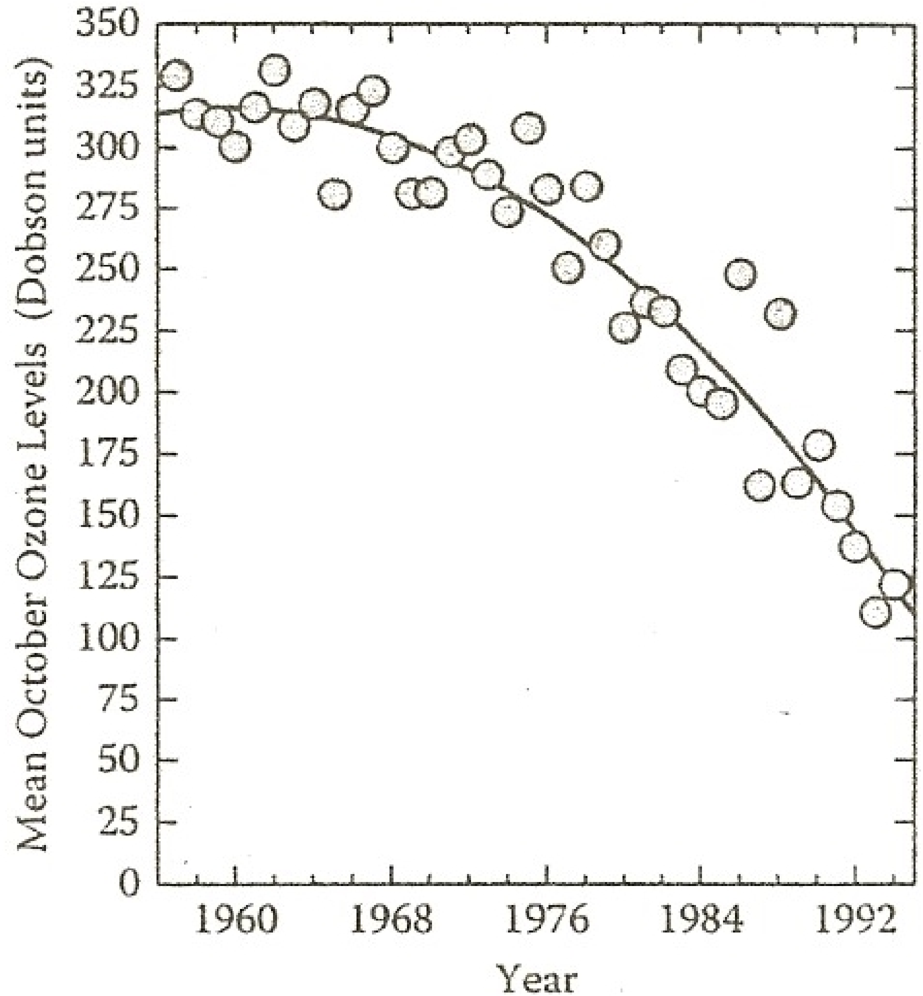


FIGURE 5.35
The decline of springtime ozone over the Antarctic over the past 33 years.

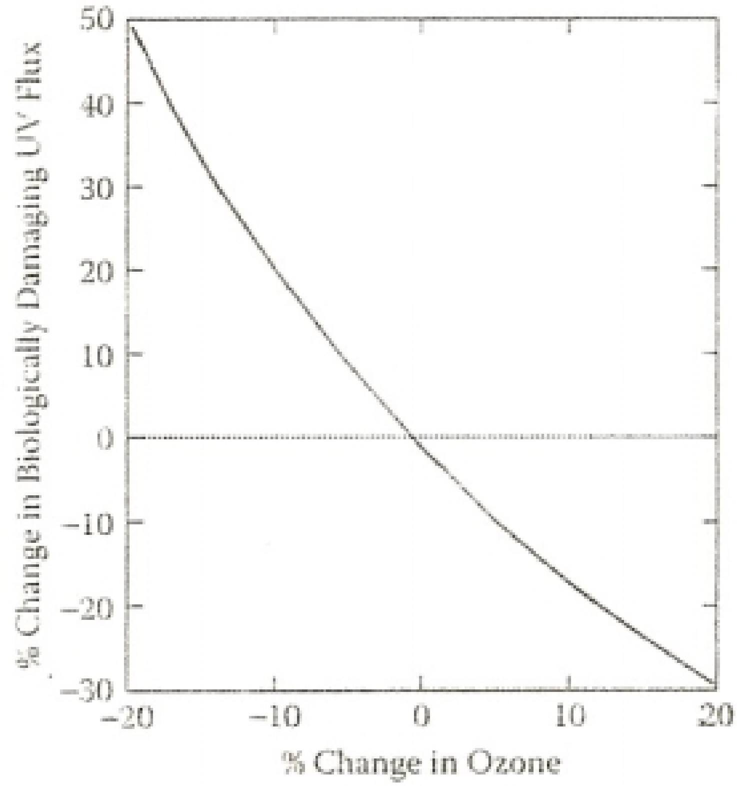


FIGURE 5.34
The changes in the biologically damaging ultraviolet (UV) flux as a function of changes in ozone levels.

SUMMARY OF OZONE CHEMISTRY

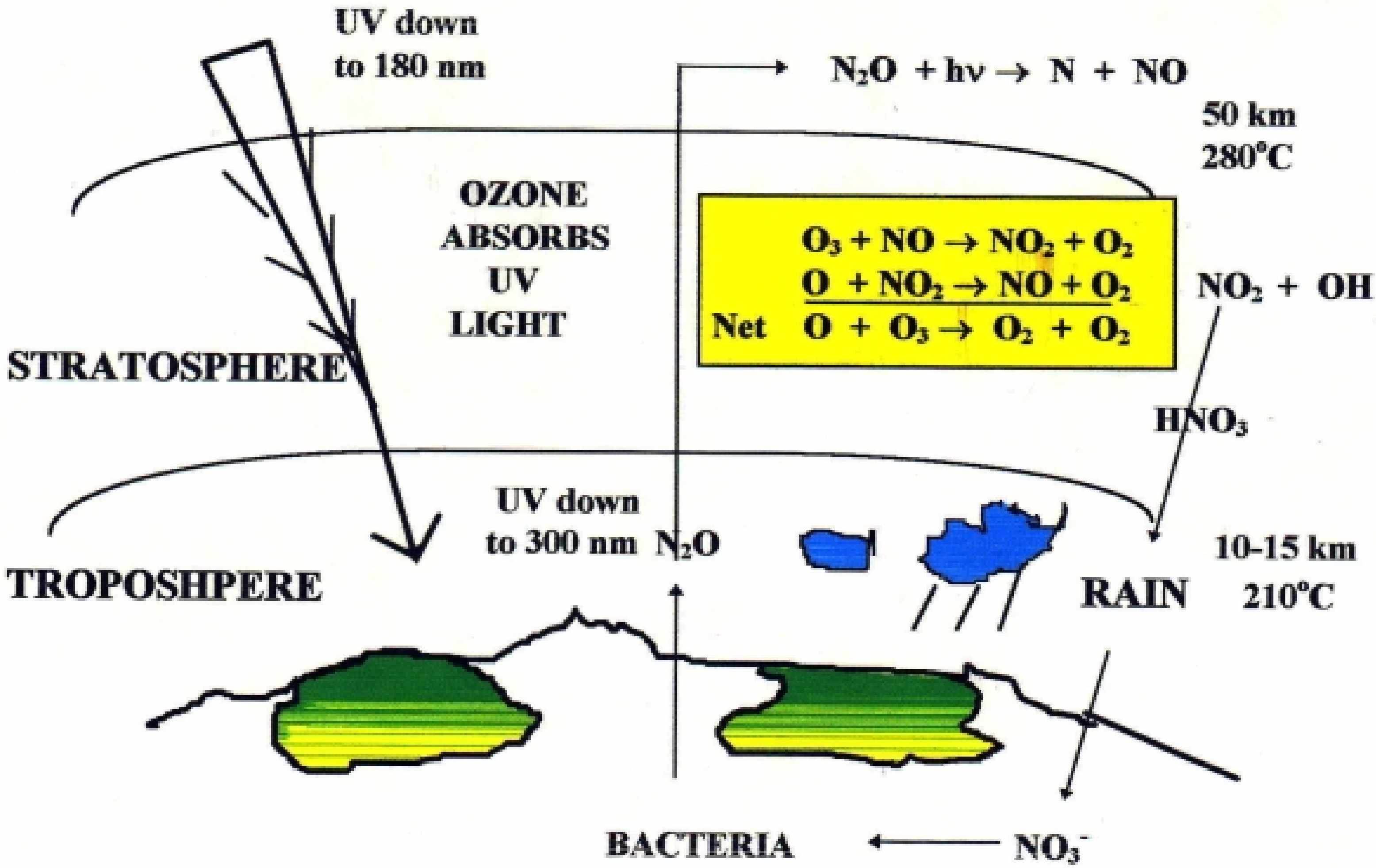
FORMATION

$O_2 + h\nu\ (< 240\ nm) \rightarrow 2\ O\cdot$

$O_2 + O\cdot \rightarrow O_3$

$O_3 + h\nu\ (240\text{-}300\ nm) \rightarrow O(^1D) + O_2$

$O(^1D) + H_2O \rightarrow 2\ OH\cdot$



UV down
to 180 nm
N₂O + hν → N + NO
50 km
280°C
OZONE
ABSORBS
UV
LIGHT
O₃ + NO → NO₂ + O₂
O + NO₂ → NO + O₂
Net O + O₃ → O₂ + O₂
NO₂ + OH
STRATOSPHERE
HNO₃
UV down
to 300 nm N₂O
10-15 km
TROPOSHPERE
RAIN
210°C
BACTERIA
NO₃⁻

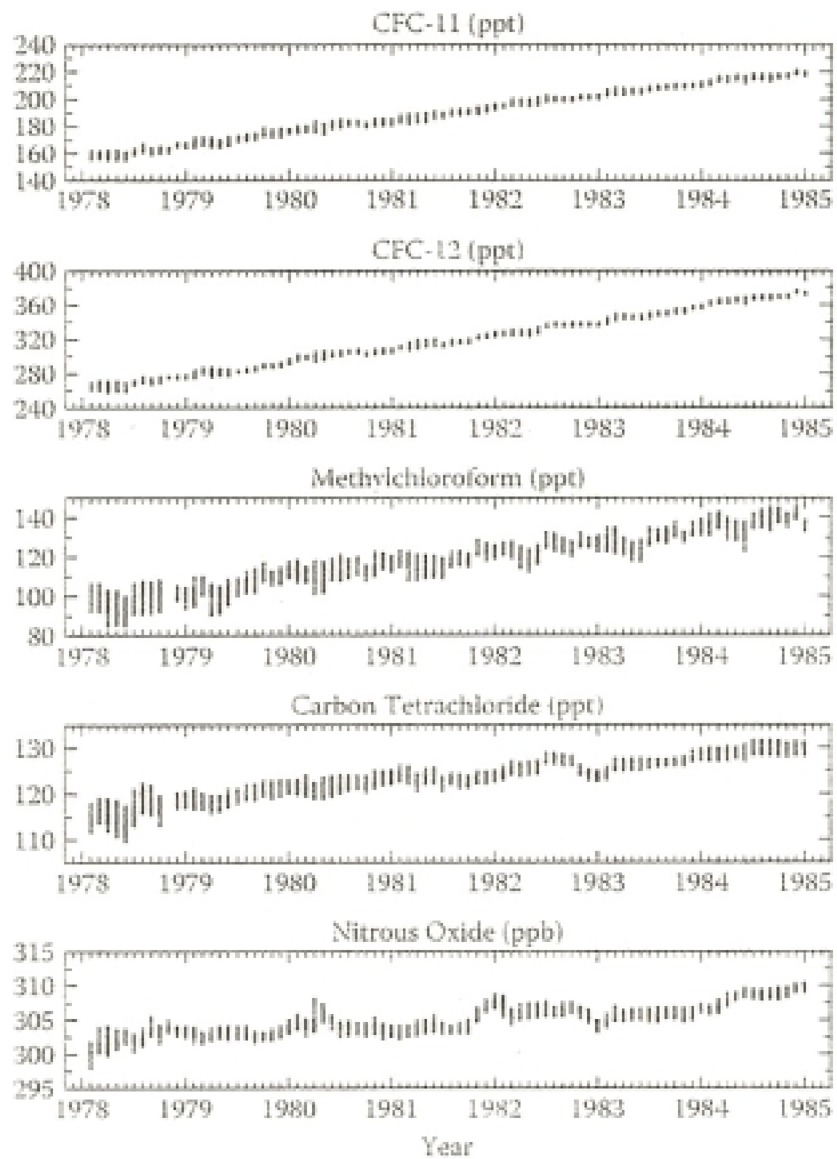


FIGURE 5.33
The increase in the concentration of chlorinated compounds and nitrogen oxide in the atmosphere.

LOSS OF OZONE IN STRATOSPHERE

Chlorofluoro Carbons (CFC's)

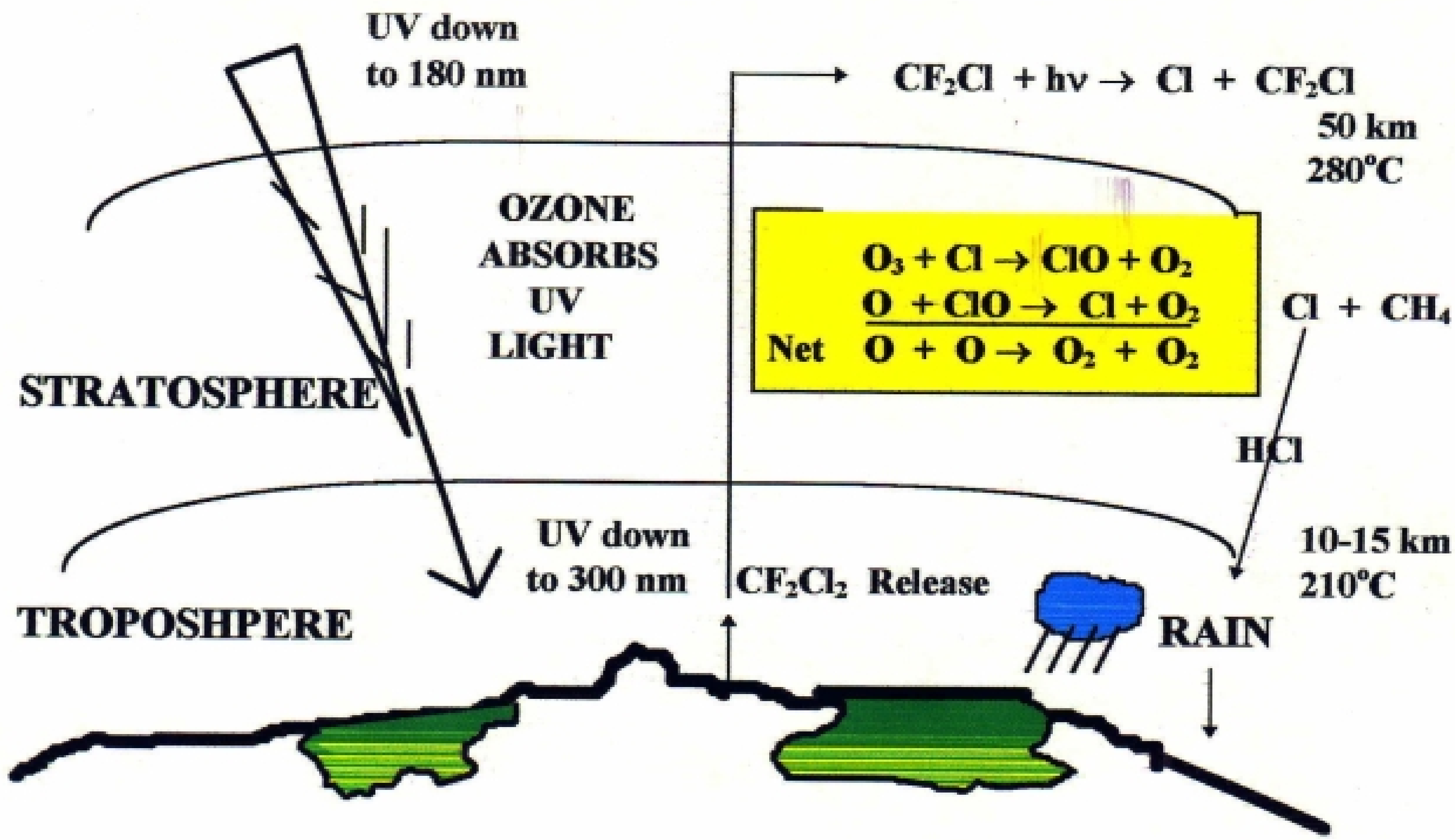
$CFCl_3$ (CFC-11) t = 75y
CF2Cl_2 (CFC-12) t = 110 y

$$CFCl_3 + h\nu \rightarrow CFCl_2 + Cl\cdot$$

$$Cl\cdot + O_3 \rightarrow ClO + O_2$$
$$ClO + O \rightarrow Cl\cdot + O_2$$
$$O + O_3 \rightarrow 2\,O_2 \quad \text{Net}$$

Scientific Support

F is found in Stratosphere
ClO appearance and O_3 loss in Antarctic


UV down
to 180 nm
CF₂Cl + hν → Cl + CF₂Cl
50 km
280°C
OZONE
ABSORBS
UV
LIGHT
O₃ + Cl → ClO + O₂
O + ClO → Cl + O₂
Net O + O → O₂ + O₂
Cl + CH₄
STRATOSPHERE
HCl
UV down
to 300 nm
CF₂Cl₂ Release
10-15 km
210°C
TROPOSHPERE
RAIN

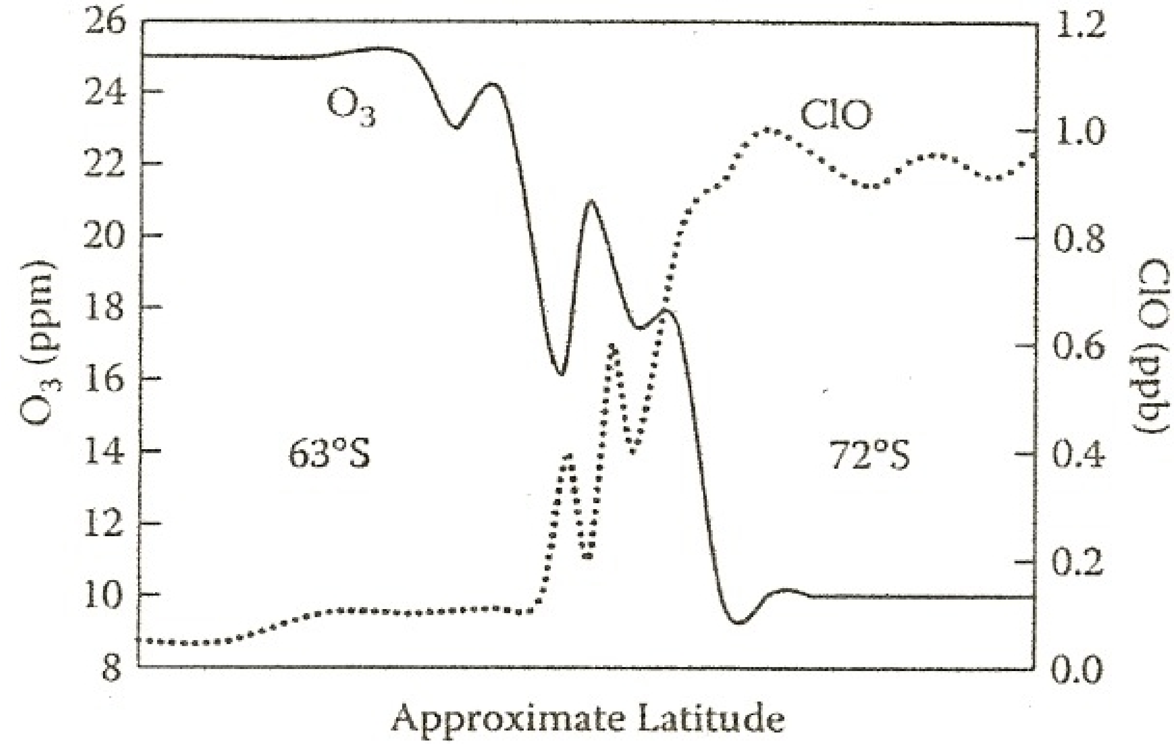


FIGURE 5.40
The relationship between O_3 and ClO in the Antarctic.

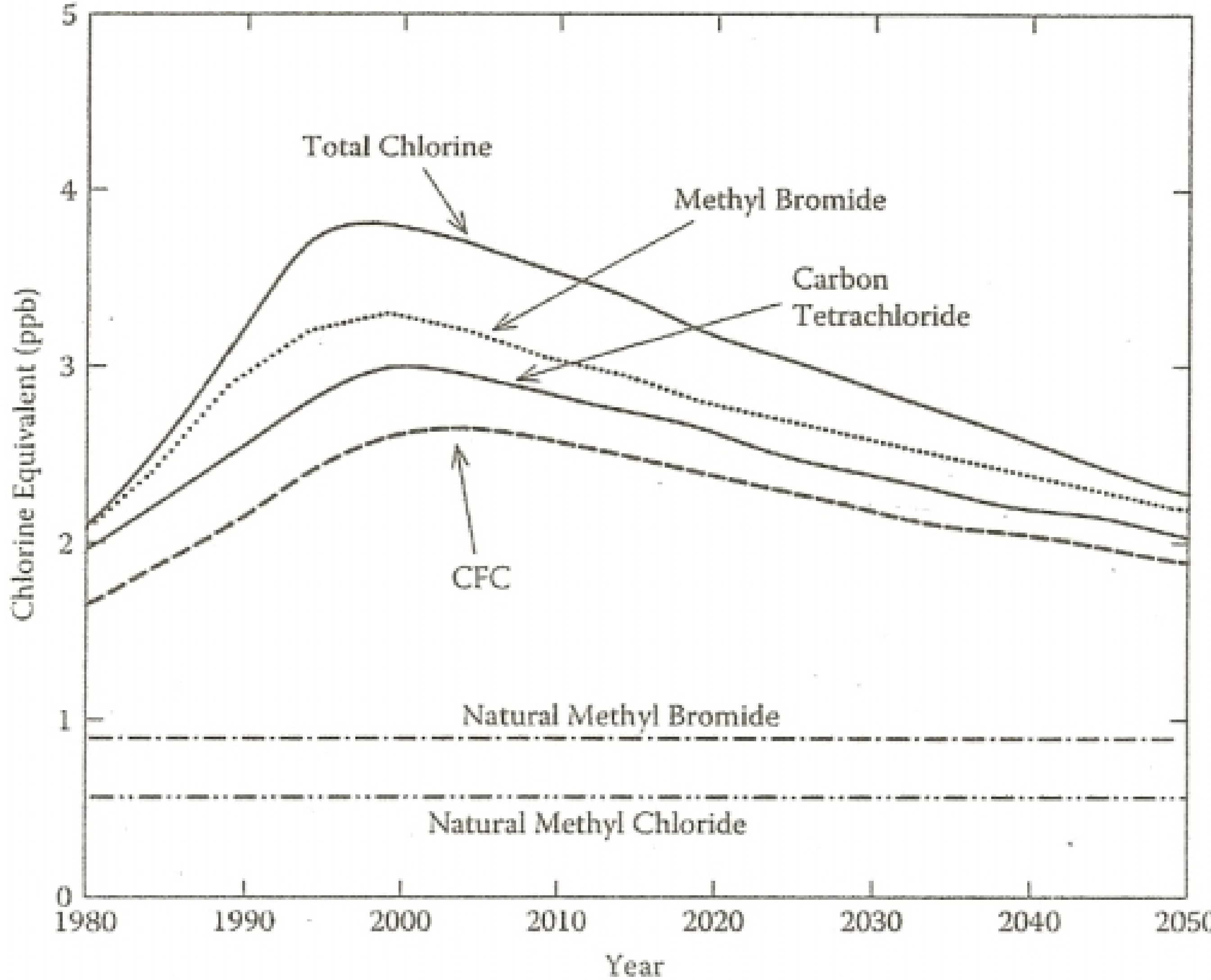
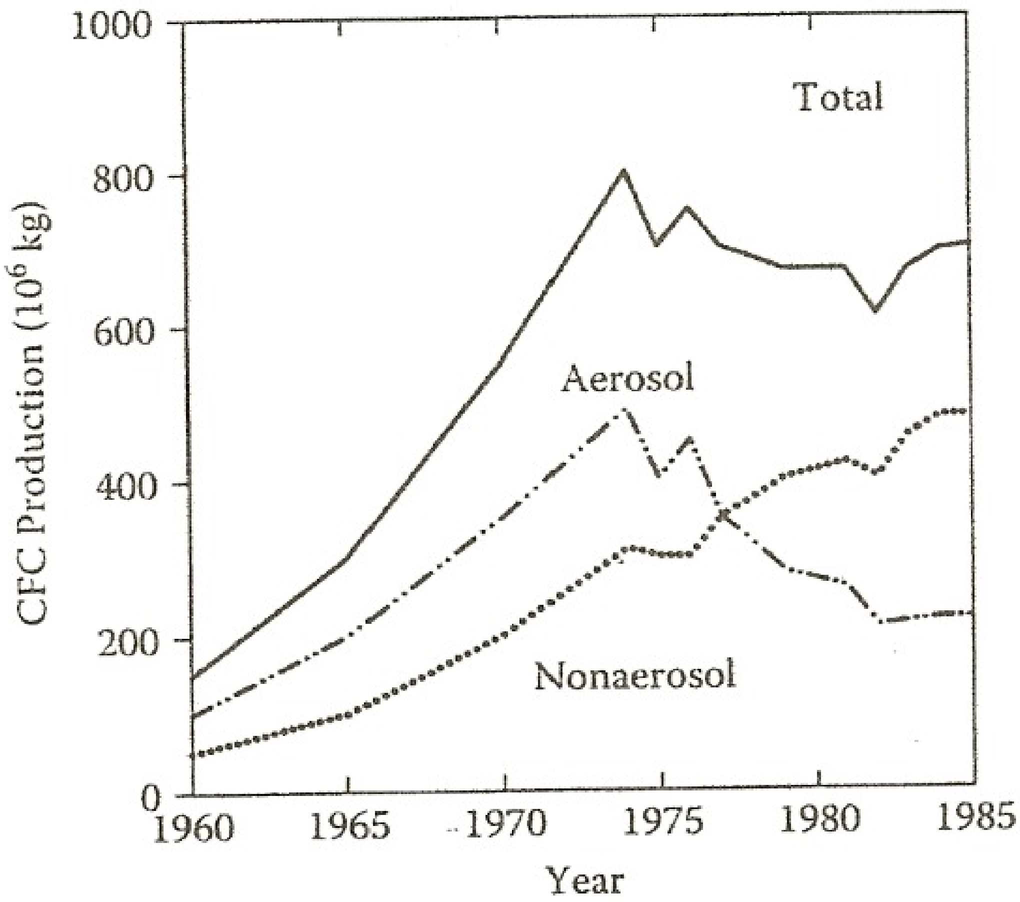


FIGURE 5.41
The levels of chlorine in the stratosphere as a result of the Montreal Protocol.



;URE 5.32
: production of CFCs from 1960 to 1985.