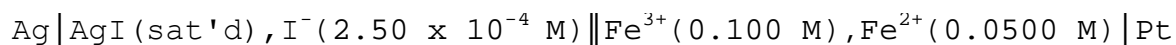
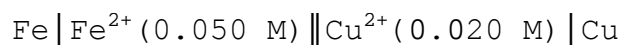


Advanced Analytical Chemistry
Homework

- 1) What is the potential of a cell made up of Zn/Zn²⁺ and Cu/Cu²⁺ half cells at 25 °C if [Zn²⁺] = 0.15 M and [Cu²⁺] = 0.25 M? Use the Nernst equation and the standard potentials given below in your calculation.
- 2) A rechargeable battery, such as a lead storage battery in a car, is an electrochemical cell that both supplies power when used and consumes power when it is recharging. Briefly (1 or 2 sentences) tell how and when this system is a galvanic cell and/or an electrolytic cell.
- 3) Calculate the potential (E) for the cells listed below. Use the Nernst equation and the list of standard potentials below.



Standard Potentials

